

Content Practice A

LESSON 2

Physical Properties

Directions: Circle the term in parentheses that correctly completes each sentence.

1. A (physical property/chemical property) is a characteristic of matter that you can observe or measure without changing the identity of the matter.
2. Water vapor is an example of a (solid/gas).
3. Mass is the amount of (matter/volume) in an object.
4. Mass is a (size-dependent/size-independent) property because its value depends on the size of a sample.
5. (Weight/Density) is the pull of gravity on matter.
6. Volume is the amount of (matter/space) something takes up.
7. Melting points and boiling points are (size-dependent/size-independent) properties of matter.
8. (Density/Weight) is the mass per unit volume of a substance.
9. Electrical conductivity is the ability of matter to (conduct/resist) an electric current.
10. Solubility is the ability of one substance to (melt/dissolve) in another.

3. **Key Concept:** What factors affect the weight of an object and not its mass?

Classified Matter

Type of Matter	Definition	Example
	1. anything that has mass and takes up space	A. atom
	2. a small particle that is a building block of matter	B. compound
	3. matter with a composition that is always the same	C. element
element	4. a substance that consists of just one type of atom	D. dissolve
	5. a substance that contains atoms of two or more elements chemically bonded together	E. heterogeneous mixture
	6. matter that can vary in composition	F. homogeneous mixture
compound	7. mixture in which the individual substances are not evenly mixed	G. matter
	8. mixture in which the individual substances are evenly mixed	H. mixture
heterogeneous mixture	9. to form a solution by mixing evenly	I. substance

homogeneous mixture		How are atoms related in elements?
		How are elements related to compounds?

2. In the space below, make a diagram that shows the general structure of an atom. Draw and label the nucleus, protons, neutrons, and electrons.

2. What can you learn about a compound from its chemical formula?

8. Compare a heterogeneous mixture and a homogeneous mixture.	
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2. Compare a solid and a liquid.	
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