

Changing Reaction Rate

1. Place one effervescent antacid tablet into a plastic cup filled with hot tap water and a second tablet into a plastic cup filled with cold water at the same time. In which cup did the fizzing last longer?
2. On a small piece of paper, crush a third tablet into a powder using a plastic spoon. Place the crushed tablet into a cup of cold water and a fourth (whole) tablet into another cup of cold water at the same time. In which cup did the fizzing last longer?
3. How does temperature affect the rate of fizzing?
4. How does crushing the reactant into a powder affect the rate of fizzing?

ANSWERS

4. Crushing the reactant into a powder caused the fizzing to happen at a faster rate.

1. The fizzing lasts longer for the tablet placed in cold water.

2. The fizzing lasts longer for the whole, uncrushed tablet.

3. Increased temperature caused the fizzing to happen at a faster rate.